**Polyatomic Ionic Compounds** (Text pgs. 91–92)

Some ions are just single atoms. Others are large molecules with several different atoms covalently bonded together. These larger ions are called **polyatomic ions**.

Eg. OH- (hydroxide), CH3COO- (acetate), NH4+ (ammonium), Cr2O72- (dichromate)



Ionic compounds that contain polyatomic ions are called **polyatomic ionic compounds**.

1. **NAMING COMPOUNDS WITH POLYATOMIC IONS (\**LOOK UP ION NAMES)***

**Metal Non-metal{use the polyatomic ion name}**

***See Table 3.10 pg. 92 IN TEXTBOOK*** *or* ***the back side of your periodic table for polyatomic ion names***

Number of Atoms List ions and determine chemical names!

Sn(OH)2

(NH4)2Cr2O7

Sr(NO3)2

Ca3(PO4)2

1. **WRITING CHEMICAL FORMULAS FOR COMPOUNDS CONTAINING POLYATOMIC IONS**

We write formulas as before, but you MUST use parentheses if there is more than one polyatomic ion required in the formula. The brackets are needed around the polyatomic ion if the subscript >1.

 Write Formula # of Atoms

 sodium permanganate

List Ions:

Balance charge:

 iron(III) sulfate

List Ions:

Balance charge:

 barium carbonate

List Ions:

Balance charge:

 magnesium bisulphite or magnesium hydrogen sulphite

List Ions:

Balance charge:

**DO pg. 91 Practice Problems**

**Section 3.2 Comprehension**

**Worksheet “BLM 1-38”**