Naming Covalent Compounds - Key

Write the formulas for the following covalent compounds:

1) Nitrogen tribromide **NBr_3**

2) hexaboron silicide **B_6Si** incorrect name hexaboron monosilicide

3) chlorine dioxide **ClO_2**

4) hydrogen iodide **HI** ionic b/c hydrogen first H^+ & I^- = HI

5) iodine pentafluoride **IF_5**

6) dinitrogen trioxide **N_2O_3**

7) ammonia (nitrogen trihydride) **NH_3** ← common name

8) phosphorus triiodide **PI_3** ← prefix name.

9) dihydrogen monoxide **H_20**

10) diphosphorous pentoxide **P_2O_5**

Write the names for the following covalent compounds:

11) P_4S_5 **tetraphosphorus pentasulfide**

12) O_2 ~~**dioxide**~~ **(oxygen)** ← proper name it is a diatomic molecule name the element.

13) SeF_6 **selenium hexafluoride**

14) Si_2Br_6 **disilicon hexabromide**

15) SCl_4 **sulfur tetrachloride**

16) CH_4 **carbon tetrahydride (methane)**

17) B_2Si **diboron silicide**

18) NF_3 **nitrogen trifluoride**

19) H_2O **dihydrogen monoxide (water)**

20) N_2O_5 **dinitrogen pentoxide**

Cl_2 = Chlorine ← not dichloride. ×