

Mixed Formula Practice Sheet

	<u>Ions combined</u>	<u>Formula</u>	<u>Chemical Name</u>
1.	potassium + oxygen	_____	_____
2.	sodium + bromine	_____	_____
3.	aluminum + iodine	_____	_____
4.	silver + sulphur	_____	_____
5.	calcium + nitrogen	_____	_____
6.	calcium + nitrate	_____	_____
7.	lithium + carbonate	_____	_____
8.	zinc + nitrate	_____	_____
9.	ammonium + hydroxide	_____	_____
10.	sodium + bicarbonate	_____	_____
11.	iron (II) + sulphur	_____	_____
12.	mercury (I) + sulphate	_____	_____
13.	nickel (III) + carbonate	_____	_____
14.	copper (II) + chlorine	_____	_____
15.	chromium (VI) + chlorate	_____	_____

- REMEMBER:**
1. The positive ion is given first in the formula
 2. The net charge of the molecule must be 0. $+2 -2 = 0$
The atoms / ions are balanced by using subscripts.
 3. A chemical formula always has the ions / subscripts in the simplest whole number ratio (i.e. reduce, if possible)
 4. When only two elements are involved, the ending is "ide."
 5. When polyatomic ions are involved, the ending is not changed. Use polyatomic ion name.
 6. When using multi-valence metals (more than one possible positive charge) make sure you include the Roman numeral in the name.

Naming Ionic Compounds Worksheet One

Give the name of the following ionic compounds:

- 1) Na_2CO_3 _____
- 2) NaOH _____
- 3) MgBr_2 _____
- 4) KCl _____
- 5) FeCl_2 _____
- 6) FeCl_3 _____
- 7) Zn(OH)_2 _____
- 8) BeSO_4 _____
- 9) CrF_2 _____
- 10) Al_2S_3 _____
- 11) PbO _____
- 12) Li_3PO_4 _____
- 13) TlI_4 _____
- 14) Co_3N_2 _____
- 15) Mg_3P_2 _____
- 16) $\text{Ga(NO}_2)_3$ _____
- 17) Ag_2SO_3 _____
- 18) NH_4OH _____
- 19) Al(CN)_3 _____
- 20) $\text{Be(CH}_3\text{COO)}_2$ _____

For the following compounds, give the formulas

- 22) sodium phosphide _____
- 23) magnesium nitrate _____
- 24) lead (II) sulfite _____
- 25) calcium phosphate _____
- 26) ammonium sulfate _____
- 27) silver cyanide _____
- 28) aluminum sulfide _____
- 29) beryllium chloride _____
- 30) copper (I) arsenide _____
- 31) iron (III) oxide _____
- 32) gallium nitride _____
- 33) iron (II) bromide _____
- 34) vanadium (V) phosphate _____
- 35) calcium oxide _____
- 36) magnesium acetate _____
- 37) aluminum sulfate _____
- 38) copper (I) carbonate _____
- 39) barium oxide _____
- 40) ammonium sulfite _____
- 41) silver bromide _____
- 42) lead (IV) nitrite _____