

key

Mixed Formula Practice Sheet

	<u>Ions combined</u>	<u>Formula</u>	<u>Chemical Name</u>
1.	potassium + oxygen	K_2O	potassium oxide
2.	sodium + bromine	$NaBr$	sodium bromide
3.	aluminum + iodine	AlI_3	aluminum iodide
4.	silver + sulphur	Ag_2S	silver sulphide
5.	calcium + nitrogen	Ca_3N_2	calcium nitride
6.	calcium + nitrate	$Ca(NO_3)_2$	calcium nitrate
7.	lithium + carbonate	Li_2CO_3	lithium carbonate
8.	zinc + nitrate	$Zn(NO_3)_2$	zinc nitrate
9.	ammonium + hydroxide	NH_4OH	ammonium hydroxide
10.	sodium + bicarbonate	$NaHCO_3$	sodium bicarbonate
11.	iron (II) + sulphur	FeS	iron(II) sulphide
12.	mercury (I) + sulphate	Hg_2SO_4	mercury(I) sulphate
13.	nickel (III) + carbonate	$Ni_2(CO_3)_3$	nickel(III) carbonate
14.	copper (II) + chlorine	$CuCl_2$	copper(II) chloride
15.	chromium (VI) + chlorate	$Cr(ClO_3)_4$	chromium(VI) chlorate

- REMEMBER:**
- The positive ion is given first in the formula
 - The net charge of the molecule must be 0. $+2 -2 = 0$
The atoms / ions are balanced by using subscripts.
 - A chemical formula always has the ions / subscripts in the simplest whole number ratio (i.e. reduce, if possible)
 - When only two elements are involved, the ending is "ide."
 - When polyatomic ions are involved, the ending is not changed. Use polyatomic ion name.
 - When using multi-valence metals (more than one possible positive charge) make sure you include the Roman numeral in the name.